

Vocabulary Check**Chapter 9****BLM 9-6****Solution Terminology****Goal**

To practice using solution terminology.

What To Do

Choose the correct term for each blank space from the list at the bottom of the page.

- The experimental evidence that NaCl in aqueous solution is a(n) _____ tells us that this salt is an electrolyte.
- $\text{Fe}^{3+}_{(\text{aq})} + 3\text{Cl}^{-}_{(\text{aq})} + 3\text{Na}^{+}_{(\text{aq})} + 3\text{OH}^{-}_{(\text{aq})} \rightarrow \text{Fe}(\text{OH})_{3(\text{s})} + 3\text{Na}^{+}_{(\text{aq})} + 3\text{Cl}^{-}_{(\text{aq})}$ is an example of a(n) _____.
- In the preceding question, the ions which do not participate in the reaction are called _____ ions.
- Since potassium chloride is a(n) _____, it exists in water solution almost completely as ions.
- Flame tests can be used to identify many _____ ions in solution.
- Spectator ions are not included in a(n) _____.
- _____ generally means that more than 1 g of solute will dissolve in 100 mL of water at room temperature.
- The bond between small ions is _____ than the bond between large ions with the same charge.
- Fluoride ions are _____ than chlorides, bromides and iodides because the fluoride ion is smaller.
- The appearance of a _____ indicates that an insoluble compound is present.
- _____ means that the solubility is less than 0.1 g per 100 mL water.
- _____ reactions can be recognized by the formation of a precipitate, gas or water.

conductor	double displacement	net ionic equation
ionic equation	metallic	soluble
insoluble	stronger	less soluble
precipitate	salt	spectator