

**Assessment****Chapter 8****BLM 8-4****Polar Versus Non-Polar****Goal**

To practice identifying different kinds of solutions and solids.

**What To Do**

In general, “like dissolves like,” so that polar solvents dissolve ionic solids and polar molecules, and non-polar solvents dissolve non-polar molecules. Alcohols, which have properties of both, tend to dissolve both types of solutes, but cannot dissolve ionic solids. Indicate which solvent the following solutes will dissolve in by checking the appropriate columns.

	Solutes	Solvents		
		Water	CCl <sub>4</sub>	Alcohol
a	NaI			
b	Br <sub>2</sub>			
c	Methanol			
d	Benzene			
e	KClO <sub>3</sub>			
f	KMnO <sub>4</sub>			
g	C <sub>6</sub> H <sub>12</sub> O <sub>6</sub>			
h	C <sub>3</sub> H <sub>8</sub>			

**Electrolyte vs. Non-Electrolytes**

Classify the following compounds as either an electrolyte or a nonelectrolyte by checking the appropriate column.

	Compound	Electrolyte	Nonelectrolyte
i	KF		
j	C <sub>12</sub> H <sub>22</sub> O <sub>11</sub>		
k	NaOH		
l	CH <sub>3</sub> OH		
m	MgCl <sub>2</sub>		
n	H <sub>2</sub> CO <sub>3</sub>		
o	C <sub>6</sub> H <sub>12</sub>		