

Reinforcement**Chapter 7****BLM 7-4****Mass/Mass Problems****Goal**

Reinforce understanding of mass-to-mass stoichiometric calculations.

What To Do

For each of the problems listed below, complete the following:

- write a balanced chemical equation
- set-up a table of data showing what is given and what is required
- solve the mass as asked for in the problem
- write a concluding statement

1. Calculate the mass of iron(III) oxide (rust) produced by the reaction of 500 g of iron with oxygen from the air.

2. What mass of precipitate should form if 2.00 g of silver nitrate in solution is reacted with excess sodium sulfide solution?

3. Determine the mass of water vapour formed when 1.00 g of butane, C_4H_{10} , is burned in a lighter.