

Reinforcement**Chapter 10****BLM 10-1**

Writing Balanced Ionic, Net Ionic, and Half-Reaction Equations

Goal

Practise writing ionic, net ionic and half-reaction equations.

Procedure

Examine the following situations. If appropriate, write balanced ionic and net ionic equations to describe the situations. If the situations are redox reactions, write oxidation and reduction half-reaction equations. Refer to the activity series for metals and solubility guidelines if necessary.

1. A solution of lead(II) nitrate is mixed with a solution of potassium iodide.

balanced ionic equation:

net ionic equation:

oxidation half-reaction:

reduction half-reaction:

2. Liquid sodium chloride is electrolyzed.

balanced ionic equation:

net ionic equation:

oxidation half-reaction:

reduction half-reaction:

3. Aqueous gold(III) chloride is stirred with a cadmium rod.

balanced ionic equation:

net ionic equation:

oxidation half-reaction:

reduction half-reaction:

4. Zinc nitrate solution is placed in a solid silver cup.

balanced ionic equation:

net ionic equation:

oxidation half-reaction:

reduction half-reaction: